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Section 8.4 - Polar Bonds and Molecules. Covalent bonds involve sharing electrons between atoms. When

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the atoms in the bond pull equally, the bonding electrons are shared equally, and the bond is nonpolar. When the atoms in the bond pull unequally, the bonding electrons are pulled closer to one atom, and the bond is polar.

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These are the vocabulary and key questions for chapter 8 of the chemistry textbook. I hope you appreciate it as much as I do, since it took me 1 1/2 hours to fill out this information. =)

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Chapter 8: Covalent Bonding. a bond angle of 109.5 degrees that results when a central

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atom forms four bonds directed toward the center of a regular tetrahedron. the mixing of several atomic orbitals to form the same total number of equivalent hybrid orbitals.

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atom or molecule: a. As you should remember, ionic compounds are solids at room temperatures that have one ion strip the electron(s) from the other elements' electron. nonpolar covalent bond. atoms in a bond share the bonding electrons equally. structural formula ...

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Covalent Bonding

Single Covalent Bonds

When only one pair of electrons is shared, such as in a hydrogen molecule, it is a single covalent bond. The shared electron pair is often referred to as the bonding pair. For a hydrogen molecule, shown in Figure 8.4, each covalently bonded atom equally

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attracts the pair of shared electrons.

Chapter 8: Covalent Bonding

Chapter 8 Covalent Bonding and Molecular Structure 8-3 There are two types of repulsive forces between the two atoms. First, the nuclei repel because they are both positively charged. Second, the electrons repel because they are both negatively charged.

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The attractive forces between the two atoms result from the

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sophisticated descriptions of chemical bonding will be discussed in chapter 9 83 lewis structures owl opening exploration 8x one of the most important tools chemists use to predict the properties of a

Chapter 8 Covalent Bonding Section 1 Molecular Compounds ...

Chapter 8: Covalent

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Bonding. Matter takes many forms in nature: In this chapter, we are going to learn to distinguish the type of compound that we have already studied, the “ionic compound” (which contains oppositely-charged particles: metal cations and non-metal anions), from a different type of compound - a “molecular compound”.

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Experimental evidence, however, indicates ...

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more strongly and gain
a slightly negative
charge in their bonds.

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